

## Article

# Competition in Coordination Assemblies: 1D-Coordination Polymer or 2D-Nets Based on Co(NCS)<sub>2</sub> and 4'-(4-methoxyphenyl)-3,2':6',3''-terpyridine

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**Abstract:** The synthesis and characterization of 4'-(4-methoxyphenyl)-3,2':6',3''-terpyridine (**2**) (IUPAC PIN 2<sup>4</sup>-(4-methoxyphenyl)-1<sup>2</sup>,2<sup>2</sup>:2<sup>6</sup>,3<sup>2</sup>-terpyridine) are described, and its coordination behaviour with cobalt(II) thiocyanate has been investigated. In a series of experiments, crystals were grown at room temperature by layering a MeOH solution of Co(NCS)<sub>2</sub> over a CHCl<sub>3</sub> solution of **2** using 1:1, 1:2 or 2:1 molar ratios of metal salt-to-ligand. Crystals harvested within 2–3 weeks proved to be the 1D-coordination polymer [Co(2)(NCS)<sub>2</sub>(MeOH)<sub>2</sub>]<sub>n</sub> and powder X-ray diffraction (PXRD) confirmed that the crystals selected for single-crystal X-ray diffraction were representative of the bulk samples. Longer crystallization times with a Co(NCS)<sub>2</sub> to **2** molar ratio of 1:1 yielded crystals of [Co(2)(NCS)<sub>2</sub>(MeOH)<sub>2</sub>]<sub>n</sub> (1D-chain) and the pseudopolymorphs [{Co(2)<sub>2</sub>(NCS)<sub>2</sub>·3MeOH}<sub>n</sub> and [{Co(2)<sub>2</sub>(NCS)<sub>2</sub>·2.2CHCl<sub>3</sub>]<sub>n</sub> ((4,4)-nets), each type of crystal originating from a different zone in the crystallization tube. PXRD for this last experiment confirmed that the dominant product in the bulk sample was the 1D-coordination polymer.

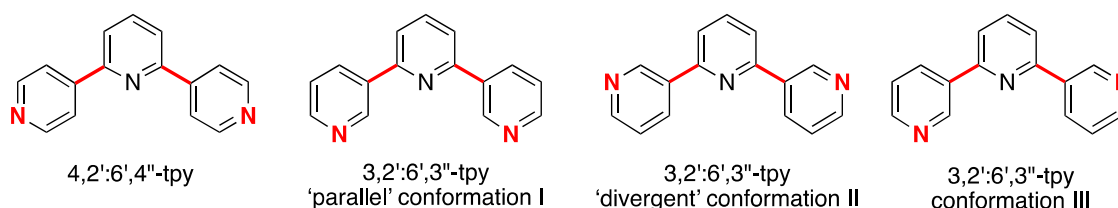
**Keywords:** 3,2':6',3''-terpyridine; coordination polymer; coordination network; cobalt(II) thiocyanate

## 1. Introduction

The use of vectorially preorganized building blocks for the rational design of 2- and 3-dimensional metal coordination architectures is a well-developed field [1,2], and leads to the tailored assembly of metal-organic frameworks (MOFs) with widespread applications [3] including gas storage [4,5], separation [6], photocatalysis [7], drug delivery [8–10], and water purification [11,12]. A key factor in controlling MOF assembly is the use of rigid organic linkers such as 4,4'-bipyridine. However, the conflict between achieving a predetermined or a serendipitous assembly is an underlying problem for the coordination chemist or crystal engineer. Shimizu has illustrated that there is scope for developing metal coordination assemblies with more flexible ligands, proposing that a ligand with “less well-defined properties does not necessarily translate to less functional materials” [13] and we have described similar self-assembly processes as examples of fuzzy logic [14].

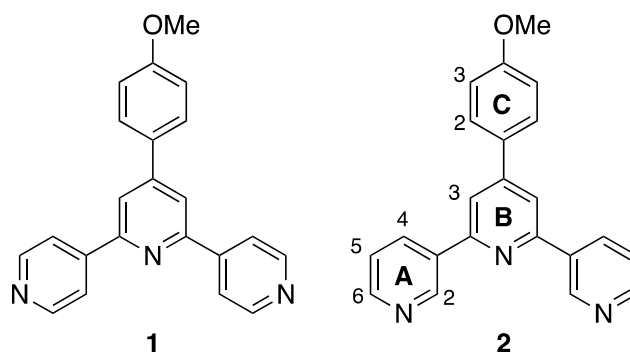
We have been focusing on metal coordination assemblies based on two isomers of terpyridine, 4,2':6',4''- and 3,2':6',3''-terpyridine (IUPAC PIN 1<sup>4</sup>,2<sup>2</sup>:2<sup>6</sup>,3<sup>4</sup>- and 1<sup>3</sup>,2<sup>2</sup>:2<sup>6</sup>,3<sup>3</sup>-terpyridine [15]) (Scheme 1) [16–19]. The nitrogen atom of the central pyridine ring is not involved in metal coordination and, therefore, these ligands are ditopic [18]. As we have previously discussed [17,18], rotation about the inter-ring C–C bonds (Scheme 1) maintains a constant vectorial relationship of the two nitrogen donors in the V-shaped 4,2':6',4''-tpy shown in Scheme 1. In contrast, the vectorial arrangement of the

two nitrogen atoms marked in red in 3,2':6',3''-tpy in Scheme 1 is dependent upon conformational changes in the ligand. Functionalization of the tpy unit in the 4'-position is easily achieved using the one-pot method of Wang and Hanan [20], and we have applied this strategy to investigate the effects on metal coordination assemblies of introducing a range of substituents into 4,2':6',4''-tpy on metal coordination assemblies. Systematic studies in which only one parameter is varied at one time provide the most valuable insight into design strategies to direct coordination architectures. However, such investigations are relatively scarce in the literature (for example, [21,22]), and analysis of data from different research groups is often complicated by use of different solvents and crystallization methods. The fact that two or more metalloassemblies (e.g., metallocycles, metallopolymers) may be close in energy is the underlying principle for the development of libraries of architectures (for example, [23–27]).



**Scheme 1.** Structures of the isomers 4,2':6',4''- and 3,2':6',3''-terpyridine (tpy), showing examples of the different conformations of 3,2':6',3''-tpy. Coordination of metal ions occurs through the outer N-donors (shown in red). See text for the significance of the C–C bonds marked in red.

We have previously investigated the coordination behaviour of 4'-(4-methoxyphenyl)-4,2':6',4''-terpyridine, **1**, (Scheme 2). A 1D-coordination polymer  $[\text{Zn}_2(\mu\text{-OAc})_4(\mathbf{1})]_n$  formed under room temperature conditions of layering an MeOH solution of  $\text{Zn}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$  over a  $\text{CHCl}_3$  solution of **1** [21]. The assembly of a 1D-chain arises from the linear disposition of the vacant axial coordination sites in a paddle-wheel  $\{\text{Zn}_2(\mu\text{-OAc})_4\}$  unit [2,28], and from the input molar ratio of  $\text{Zn}(\text{OAc})_2$  to **1** which was 2:1 [21]. Using similar layering conditions at ambient temperatures, reaction of **1** with cadmium(II) nitrate led to the assembly of a ladder  $[\{\text{Cd}_2(\mathbf{1})_3(\text{NO}_3)_4\} \cdot \text{CHCl}_3 \cdot \text{MeOH}]_n$  in which the ligand **1** acts as a linker, forming both the “rungs” and “rails” of the ladder. The 2:3 Cd to **1** ratio in the structure contrasted with the input molar ratio of 3:1, but analysis of other crystals from separate crystallization tubes confirmed this stoichiometry for the crystalline coordination polymer [29]. In directing assemblies, cobalt(II) thiocyanate is a potential 4-connecting node, i.e., octahedral coordination with two sites occupied by thiocyanato ligands. Therefore, the observed assembly of a (4,4) net from a combination of **1** and  $\text{Co}(\text{NCS})_2$  is both readily understood and expected. The 2D-net  $[\{\text{Co}_2(\text{NCS})_4(\mathbf{1})_4\} \cdot 2\text{CHCl}_3 \cdot 1.5\text{MeOH}]_n$  was obtained under conditions of layering an MeOH solution of  $\text{Co}(\text{NCS})_2$  over a  $\text{CHCl}_3$  solution of **1**, even though the molar input ratio of Co:**1** was 1:3 [30]. A similar (4,4) net,  $[\{\text{Co}(\mathbf{1})_2(\text{NCS})_2\} \cdot 2\text{MeCN} \cdot 2\text{DMF}]_n$  (DMF = dimethylformamide), has been reported by Konar and co-workers, although in this case, hydrothermal conditions were used with an input molar ratio of Co to **1** of 2:1 [31]. Therefore, in both cases, irrespective of crystallization method, there is a preference for a (4,4) net with a Co to **1** ratio of 1:2. An analogous (4,4) net is found in  $[\{\text{Fe}(\mathbf{1})_2(\text{NCS})_2\} \cdot 2\text{MeCN}]_n$ , again formed hydrothermal conditions [31]. Based on a search of the Cambridge Structural Database (CSD, v. 5.40 with February 2019 updates [32], using Conquest v. 2.0.1 [33]), the only other reports of structurally characterized coordination networks involving ligand **1** describe reactions with copper(I) iodide under hydrothermal synthetic conditions to give a series of related (4,4) nets with 4-connecting  $\{\text{Cu}_2\text{I}_2\}$  nodes [34,35]. The intriguing difference between these 2D-networks and those based on  $\text{Co}(\text{NCS})_2$  is that the copper(I)-containing assemblies are interpenetrating, leading to porous 3D-frameworks which are thermally robust and can undergo reversible solvent exchange [34].



**Scheme 2.** Structures of ligands **1** and **2** with atom numbering scheme for the NMR assignments of **2**.

Noting the observations discussed above, we decided to carry out an investigation of the reaction of ligand **2** (Scheme 2) with  $\text{Co}(\text{NCS})_2$ , with the aim of comparing the coordination of the isomeric ligands **1** and **2** with  $\text{Co}(\text{NCS})_2$ , and to gain a better understanding of the effects of varying the metal to ligand ratio in the crystallization tubes.

## 2. Materials and Methods

### 2.1. General

$^1\text{H}$  and  $^{13}\text{C}\{^1\text{H}\}$  NMR spectra were recorded on a Bruker Avance iii-500 spectrometer (Bruker BioSpin AG, Fällanden, Switzerland) at 298 K. The  $^1\text{H}$  and  $^{13}\text{C}$  NMR chemical shifts were referenced with respect to residual solvent peaks ( $\delta$  TMS = 0). Electrospray ionization (ESI) mass spectra were recorded using a Shimadzu LCMS-2020 instrument (Shimadzu Schweiz GmbH, Reinach, Switzerland) samples were introduced as 200–800  $\mu\text{M}$  solutions in MeCN with the addition of formic acid. PerkinElmer UATR Two (Perkin Elmer, Schwerzenbach, Switzerland) and Cary-5000 (Agilent Technologies Inc., Santa Clara, CA, USA) spectrometers were used to record FT-infrared (IR) and absorption spectra, respectively.

3-Acetylpyridine was purchased from Acros Organics (Chemie Brunschwig AG, Basel, Switzerland), and 4-methoxybenzaldehyde and  $\text{Co}(\text{NCS})_2$  were bought from Sigma Aldrich (Steinheim, Germany). All chemicals were used as received.

### 2.2. Synthesis of **2**

4-Methoxybenzaldehyde (1.36 g, 1.22 mL, 10.0 mmol) was dissolved in EtOH (50 mL), and then 3-acetylpyridine (2.42 g, 2.20 mL, 20.0 mmol) and crushed KOH (1.12 g, 20.0 mmol) were added, upon which the solution changed from colourless to yellow. Aqueous  $\text{NH}_3$  (32%, 38.5 mL) was slowly added to the reaction mixture. This was stirred at room temperature overnight. The solid that formed was collected by filtration, washed with water ( $3 \times 10$  mL), EtOH ( $3 \times 10$  mL), recrystallized from EtOH and dried in vacuo. Compound **2** was isolated as a white powder (0.888 g, 2.62 mmol, 26.2%). M.p. = 157 °C.  $^1\text{H}$  NMR (500 MHz,  $\text{CDCl}_3$ ):  $\delta/\text{ppm}$  = 9.38 (d,  $J$  = 2.1 Hz, 2H,  $\text{H}^{\text{A}2}$ ), 8.71 (dd,  $J$  = 4.8, 1.5 Hz, 2H,  $\text{H}^{\text{A}6}$ ), 8.55 (m, 2H,  $\text{H}^{\text{A}4}$ ), 7.94 (s, 2H,  $\text{H}^{\text{B}3}$ ), 7.72 (m, 2H,  $\text{H}^{\text{C}2}$ ), 7.50 (m, 2H,  $\text{H}^{\text{A}5}$ ), 7.08 (m, 2H,  $\text{H}^{\text{C}3}$ ), 3.90 (s, 3H,  $\text{H}^{\text{a}}$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (500 MHz,  $\text{CDCl}_3$ ):  $\delta/\text{ppm}$  = 161.1 ( $\text{C}^{\text{C}4}$ ), 155.2 ( $\text{C}^{\text{A}3}$ ), 150.7 ( $\text{C}^{\text{B}4}$ ), 149.6 ( $\text{C}^{\text{A}6}$ ), 147.9 ( $\text{C}^{\text{A}2}$ ), 135.3 ( $\text{C}^{\text{A}4}$ ), 135.2 ( $\text{C}^{\text{B}2}$ ), 130.4 ( $\text{C}^{\text{C}1}$ ), 128.5 ( $\text{C}^{\text{C}2}$ ), 124.0 ( $\text{C}^{\text{A}5}$ ), 117.5 ( $\text{C}^{\text{B}3}$ ), 114.9 ( $\text{C}^{\text{C}3}$ ), 55.6 ( $\text{C}^{\text{a}}$ ). UV–VIS ( $\text{CH}_3\text{CN}$ ,  $5.0 \times 10^{-5}$  mol  $\text{dm}^{-3}$ )  $\lambda/\text{nm}$  228 ( $\epsilon/\text{dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$  24,400), 274 (30,000). ESI-MS  $m/z$  340.10 [ $\text{M} + \text{H}$ ] $^+$  (calc. 340.14). Found C 77.74, H 5.05, N 12.36; required for  $\text{C}_{22}\text{H}_{17}\text{N}_3\text{O}$ : C 77.86, H 5.05, N 12.38.

### 2.3. $[\text{Co}(\text{2})(\text{NCS})_2(\text{MeOH})_2]_n$ , $[\{\text{Co}(\text{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$ , and $[\{\text{Co}(\text{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$

In Experiment I, a solution of  $\text{Co}(\text{NCS})_2$  (5.3 mg, 0.030 mmol) in MeOH (6 mL) was layered over a  $\text{CHCl}_3$  solution (4 mL) of **2** (10.2 mg, 0.030 mmol). Pink block-like crystals visible to the eye were first obtained after 17 days and a single crystal was selected for X-ray diffraction after

another 4 weeks. This proved to be  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ . Cell dimension checks carried out on two other block-shaped crystals from the crystallization tube confirmed consistent cell dimensions. The remaining crystals in the tube were washed with MeOH and then  $\text{CHCl}_3$ , dried under vacuum and were analysed by powder X-ray diffraction (PXRD).

In Experiment II, a solution of  $\text{Co}(\text{NCS})_2$  (5.3 mg, 0.030 mmol) in MeOH (6 mL) was layered over a  $\text{CHCl}_3$  solution (4 mL) of  $\mathbf{2}$  (10.2 mg, 0.030 mmol). Pink crystals visible to the eye started growing after one week and a single crystal selected for X-ray diffraction after a further week was found to be  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$ . A second crystal was removed from the bottom of the tube after a further 5 weeks and single-crystal X-ray diffraction confirmed this to be  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$ . A cell dimension check was carried out on one of the block-like crystals growing on the wall of the tube confirmed it to be  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ . The remaining crystals were washed with MeOH and  $\text{CHCl}_3$ , dried under vacuum and were analysed by PXRD.

In Experiment III, a solution of  $\text{Co}(\text{NCS})_2$  (10.5 mg, 0.0600 mmol) in MeOH (6 mL) was layered over a  $\text{CHCl}_3$  solution (4 mL) of  $\mathbf{2}$  (10.2 mg, 0.030 mmol). Pink blocks visible to the eye started growing after 17 days and the tube was left for 3 months at room temperature. A cell dimension check on one crystal from this tube was consistent with  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ . The remaining crystals in the tube were washed with MeOH and  $\text{CHCl}_3$ , dried under vacuum and were analysed by PXRD.

In Experiment IV, a solution of  $\text{Co}(\text{NCS})_2$  (2.6 mg, 0.015 mmol) in MeOH (5 mL) was layered over a  $\text{CHCl}_3$  solution (5 mL) of  $\mathbf{2}$  (10.2 mg, 0.030 mmol). A few pink plates visible to the eye started growing after 15 days at room temperature. A cell dimension check on one crystal from this tube was consistent with  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ . Insufficient amount of material was available for PXRD.

Bulk reactions were carried out with 1:1 and 1:2  $\text{Co}(\text{NCS})_2$  to  $\mathbf{2}$  molar ratios. Solid  $\text{Co}(\text{NCS})_2$  (51.7 mg, 0.295 mmol) was added to a colourless solution of  $\mathbf{2}$  (100 mg, 0.295 mmol) in MeOH (40 mL). Immediate precipitation of a pale pink solid was observed and the mixture was left stirring at room temperature for 15 h. The product was separated by centrifuge followed by decantation of the supernatant liquid. The solid was washed with MeOH ( $2 \times 10$  mL) and with  $\text{CHCl}_3$  (10 mL) to remove unreacted reagents and was then dried under vacuum. The procedure for the 1:2 reaction was the same, starting with  $\text{Co}(\text{NCS})_2$  (25.7 mg, 0.147 mmol) and  $\mathbf{2}$  (100 mg, 0.295 mmol). The two solid products were investigated by PXRD.

#### 2.4. Crystallography

Single crystal data for  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$  were collected on a Bruker APEX-II diffractometer with data reduction, solution and refinement using the programs APEX [36], ShelXT [37], Olex2 [38] and ShelXL v. 2014/7 [39]. For  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ , single crystal data were collected on a STOE StadiVari diffractometer equipped with a Pilatus300K detector and with a Metaljet D2 source; the structure was solved using Superflip [40,41] and Olex2 [38]. The model was refined with ShelXL v. 2014/7 [39]. Structure analysis used Mercury CSD v. 4.1.0 [42,43].

SQUEEZE [44] was used in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$ . The electrons removed equated to 3 molecules of MeOH per cobalt atom.

Powder X-Ray Diffraction (PXRD) patterns were collected at room temperature in the Debye-Scherrer geometry using a Bruker AXS D8 Advance diffractometer (Bruker AXS GmbH, Karlsruhe, Germany) equipped with a Ni-filtered  $\text{Cu K}\alpha$  radiation and a 1D LynxEye PSD detector. The reflections of a main phase were indexed with an orthorhombic cell in the space group  $Pbca$  (No 61). The Rietveld refinement analysis [45] of the diffraction patterns was performed with the package FULLPROF SUITE [46,47] (version March-2019) using a previously determined instrument resolution function (based on the small line width polycrystalline sample  $\text{Na}_2\text{Ca}_3\text{Al}_2\text{F}_{14}$  measurements [48,49]). The structural model was taken from the single-crystal X-Ray diffraction refinement. Refined parameters were: scale factor, zero displacement, lattice parameters, Co and S atomic positions, and peak shapes as a Thompson–Cox–Hastings pseudo-Voigt function. Due to the powdered crystals' habit, a preferred orientation as a March–Dollase multi-axial phenomenological model was implemented in the analysis.

### 2.5. $[\text{Co}(\text{2})(\text{NCS})_2(\text{MeOH})_2]_n$

$\text{C}_{26}\text{H}_{25}\text{CoN}_5\text{O}_3\text{S}_2$ ,  $M_r = 578.56$ , pink block, orthorhombic, space group  $Pbca$ ,  $a = 14.0281(3)$ ,  $b = 14.3649(3)$ ,  $c = 26.0077(6)$  Å,  $V = 5240.9(2)$  Å<sup>3</sup>,  $D_c = 1.467$  g cm<sup>−3</sup>,  $T = 130$  K,  $Z = 8$ ,  $Z' = 1$ ,  $\mu(\text{GaK}\alpha) = 4.713$  mm<sup>−1</sup>. Total 98166 reflections, 5359 unique ( $R_{\text{int}} = 0.3579$ ). Refinement of 3094 reflections (343 parameters) with  $I > 2\sigma(I)$  converged at final  $R_1 = 0.0639$  ( $R_1$  all data = 0.1341),  $wR_2 = 0.1376$  ( $wR_2$  all data = 0.1728),  $\text{gof} = 1.006$ . CCDC 1922228.

### 2.6. $[\{\text{Co}(\text{2})(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$

$\text{C}_{49}\text{H}_{46}\text{CoN}_8\text{O}_5\text{S}_2$ ,  $M_r = 949.99$ , pink plate, monoclinic, space group  $P2_1/n$ ,  $a = 13.2570(6)$ ,  $b = 12.1980(5)$ ,  $c = 15.0937(6)$  Å,  $\beta = 98.675(2)$ ,  $V = 2412.86(18)$  Å<sup>3</sup>,  $D_c = 1.308$  g cm<sup>−3</sup>,  $T = 130$  K,  $Z = 2$ ,  $Z' = 0.5$ ,  $\mu(\text{CuK}\alpha) = 4.069$  mm<sup>−1</sup>. Total 15819 reflections, 4303 unique ( $R_{\text{int}} = 0.0336$ ). Refinement of 3963 reflections (269 parameters) with  $I > 2\sigma(I)$  converged at final  $R_1 = 0.0386$  ( $R_1$  all data = 0.0420),  $wR_2 = 0.1092$  ( $wR_2$  all data = 0.1118),  $\text{gof} = 1.062$ . CCDC 1922226.

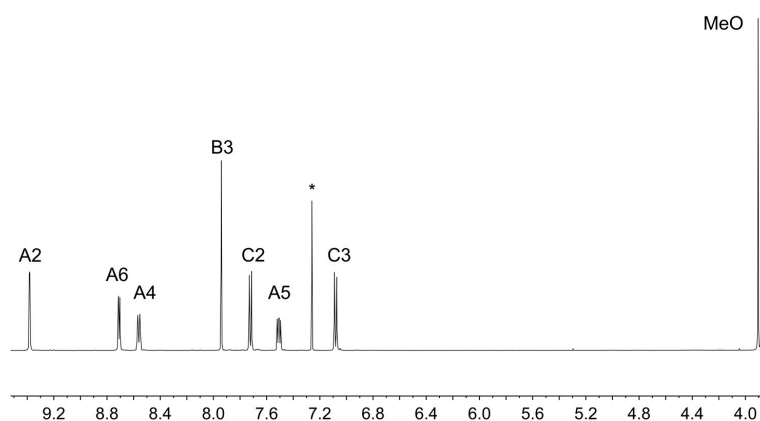
### 2.7. $[\{\text{Co}(\text{2})(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$

$\text{C}_{48.2}\text{H}_{36.6}\text{Cl}_{6.6}\text{CoN}_8\text{O}_2\text{S}_2$ ,  $M_r = 1116.87$ , pink block, monoclinic, space group  $P2_1/c$ ,  $a = 11.0555(8)$ ,  $b = 13.2414(10)$ ,  $c = 18.5561(13)$  Å,  $\beta = 96.956(3)$ ,  $V = 2696.4(3)$  Å<sup>3</sup>,  $D_c = 1.376$  g cm<sup>−3</sup>,  $T = 130$  K,  $Z = 2$ ,  $Z' = 0.5$ ,  $\mu(\text{CuK}\alpha) = 6.592$  mm<sup>−1</sup>. Total 23049 reflections, 4886 unique ( $R_{\text{int}} = 0.0313$ ). Refinement of 4652 reflections (341 parameters) with  $I > 2\sigma(I)$  converged at final  $R_1 = 0.1177$  ( $R_1$  all data = 0.1199),  $wR_2 = 0.4012$  ( $wR_2$  all data = 0.4131),  $\text{gof} = 1.096$ . CCDC 1922227.

## 3. Results and Discussion

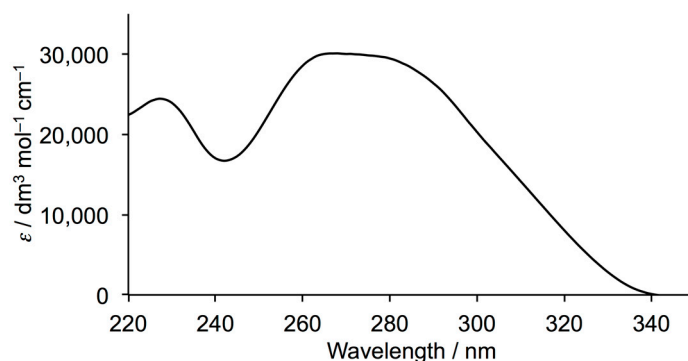
### 3.1. Ligand Synthesis and Characterization

Compound **2** was prepared by Hanan's [20] one-pot synthesis from 3-acetylpyridine and 4-methoxybenzaldehyde in the presence of base, followed by the addition of aqueous ammonia to give **2** in 26.2% yield after purification. A single peak in the electrospray mass spectrum at  $m/z = 340.10$  (Figure S1, see Supporting Information) corresponded to  $[\text{M} + \text{H}]^+$ . The <sup>1</sup>H and <sup>13</sup>C{<sup>1</sup>H} NMR spectra were assigned using 2D-methods. Figure 1 shows the <sup>1</sup>H NMR spectrum and the <sup>13</sup>C{<sup>1</sup>H} NMR, NOESY, HMQC and HMBC spectra are shown in Figures S2–S5 (see Scheme 2 for atom labelling). The spectra are consistent with a time-averaged C<sub>2</sub>-symmetric ligand and protons H<sup>C2</sup> and H<sup>C3</sup> were distinguished by the NOESY crosspeak between H<sup>B3</sup> and H<sup>C2</sup> (Figure S3). The solution absorption spectrum of **2** (Figure 2) exhibits intense absorptions arising from spin-allowed  $\pi^* \leftarrow n$  and  $\pi^* \leftarrow \pi$  transitions, and is comparable to the absorption spectra of the analogous ligands bearing ethoxy or *n*-butoxy substituents [50].



**Figure 1.** 500 <sup>1</sup>H MHz NMR spectrum of **2** ( $\text{CDCl}_3$ , 298 K). \* = residual  $\text{CHCl}_3$ . See Scheme 2 for atom labelling.





**Figure 2.** Solution absorption spectrum of compound **2** (MeCN,  $5 \times 10^{-5}$  mol dm $^{-3}$ ).

### 3.2. Reactions of Compound **2** with Co(NCS) $_2$

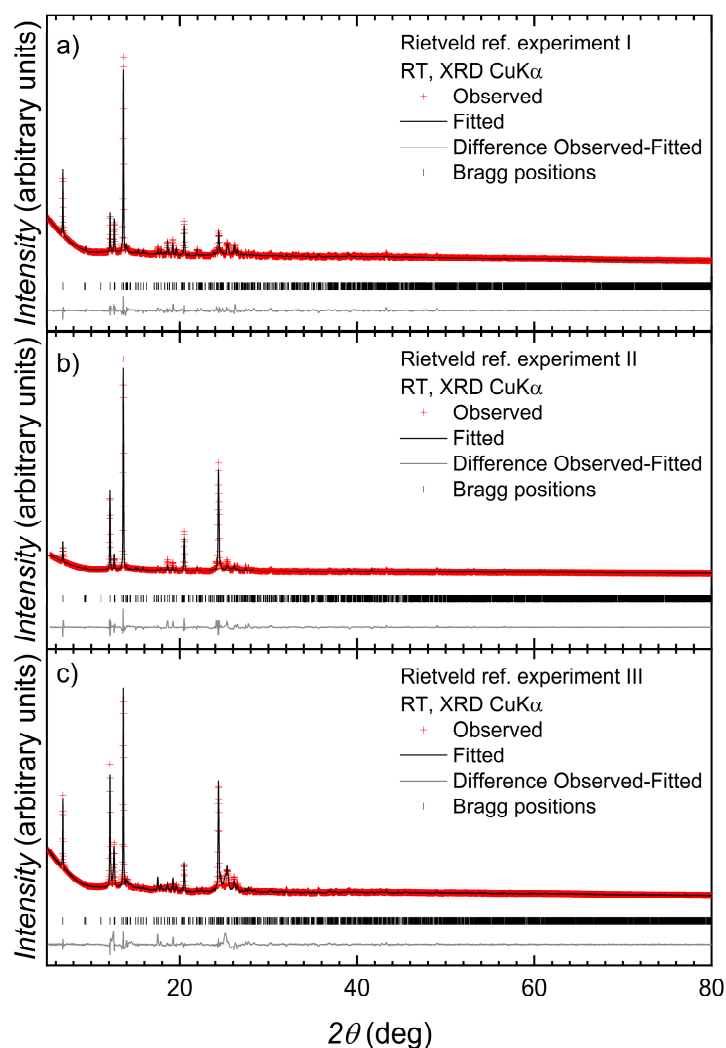
Four tubes, each containing a MeOH solution of Co(NCS) $_2$  layered over a CHCl $_3$  solution of **2**, were left standing at room temperature (ca. 295 K). In two of the tubes (Experiments I and II), identical conditions were used with a molar ratio of Co(NCS) $_2$ :**2** of 1:1. In Experiments III and IV, the same crystallization conditions were used as in Experiments I and II but with molar ratios of Co(NCS) $_2$ :**2** of 2:1 and 1:2, respectively. The experiments are summarized in Scheme 3 and details of the rate of crystal growth and the selection of crystals are given in the Materials and Methods Section 2.3.

Experiment	I	II	III	IV
Molar ratio Co: <b>2</b>	1:1	1:1	2:1	1:2
Times for harvesting crystals	17 d, 4 wk	2 wk, 7 wk	3 mth	15 d

**Scheme 3.** Summary of Experiments I–IV.

In Experiments I and III, pink block-like single crystals of [Co(**2**)(NCS) $_2$ (MeOH) $_2$ ] $_n$  were obtained, while in Tube II, three different crystal types were found in different zones in the tube. Blocks of [Co(**2**)(NCS) $_2$ (MeOH) $_2$ ] $_n$  grew on the walls of the tube. Pink block-like crystals of [Co(**2**) $_2$ (NCS) $_2$ ·2.2CHCl $_3$ ] $_n$  formed at the bottom of the tube and it is relevant that initially, the CHCl $_3$  solution of the ligand comprised the lower layer in the crystallization tube. Higher up on the walls of the tube, pink plates of [Co(**2**) $_2$ (NCS) $_2$ ·3MeOH] $_n$  were obtained, consistent with growth in a zone where the concentration of methanol exceeds that of chloroform. As described below, [Co(**2**) $_2$ (NCS) $_2$ ·2.2CHCl $_3$ ] $_n$  and [Co(**2**) $_2$ (NCS) $_2$ ·3MeOH] $_n$  are pseudopolymorphs, differing only in the inclusion of different amounts of different lattice solvent [51,52].

The PXRD patterns of the bulk material from Experiments I and III are displayed in Figures S6 and S7 and a comparison with the pattern predicted from the single crystal structure of  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$  (see below) indicates that the single crystal selected was representative of all single crystals that grew from the 1:1 and 2:1, respectively, reactions of  $\text{Co}(\text{NCS})_2$  with **2**. A Rietveld refinement of the diffraction patterns, using the structural model from the single crystal refinement, was carried out using FULLPROF SUITE [46,47]. The refinement confirmed, that the bulk sample from each of Experiments I and III is  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$  (Figure 3a,c). Preferred orientations in this material correspond to the (001) and (010) lattice planes. A bulk reaction with a 1:1 molar ratio of reactants was also carried out (see Materials and Methods Section). Although the sample was of low crystallinity, peaks at low values of  $2\theta$  (in particular that at  $2\theta = 6.82^\circ$ ) in the PXRD pattern (Figure S8) were consistent with  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ .



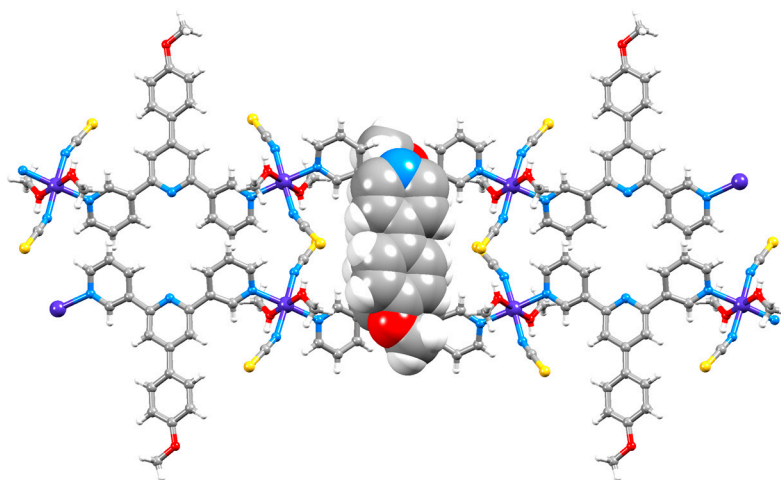
**Figure 3.** Laboratory X-ray diffraction ( $\text{CuK}\alpha$  radiation) pattern (red crosses) for Experiments (a) I, (b) II, and (c) III (space group  $Pbca$ ) at room temperature. The black line corresponds to the best fit from the Rietveld refinement analysis. Lower vertical marks denote the Bragg peak positions. The bottom, grey line represents the difference between experimental and calculated points.

Figure S9 shows a comparison of the PXRD pattern for the bulk material from Experiment II with those predicted from the single crystal structures of  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ ,  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$ . The Rietveld refinement confirmed that the bulk sample represents the 1D coordination polymer (Figure 3b). Very small amounts of other

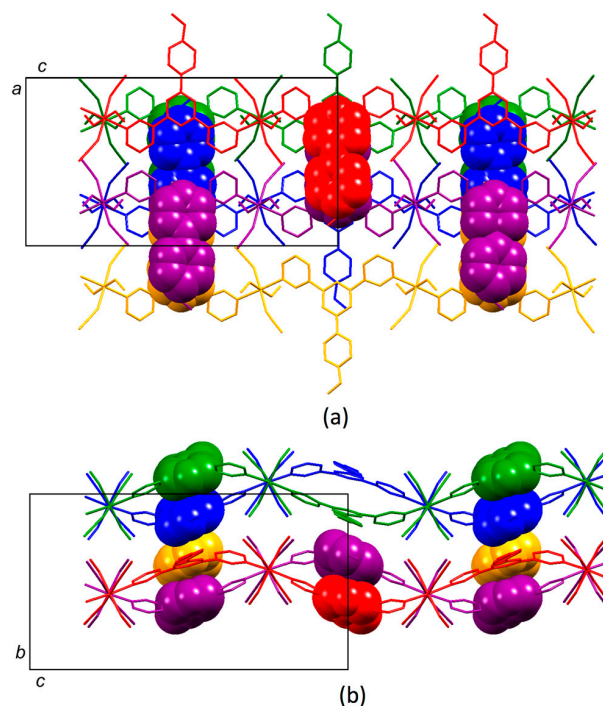




stacked units in Figure 6. The alternate stacking of the purple-coloured chain in Figure 6 with red- and orange-coloured chains leads to extended  $\pi$ -stacking which interlock the chains through the lattice.



**Figure 5.** Packing of chains in  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ . View down the  $b$ -axis showing head-to-tail  $\pi$ -stacking between 4-methoxyphenylpyridine units of pairs of ligands **2**.

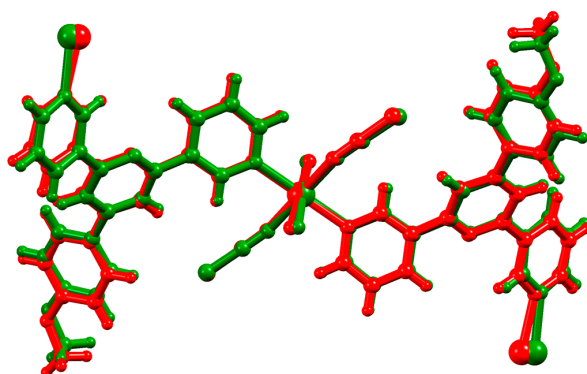


**Figure 6.** Packing of chains in  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$ .  $\pi$ -Stacking between 4-methoxyphenylpyridine units in four adjacent chains (each depicted in a different colour) is shown, with the same part of the lattice viewed down (a) the  $b$ -axis and (b) the  $a$ -axis.

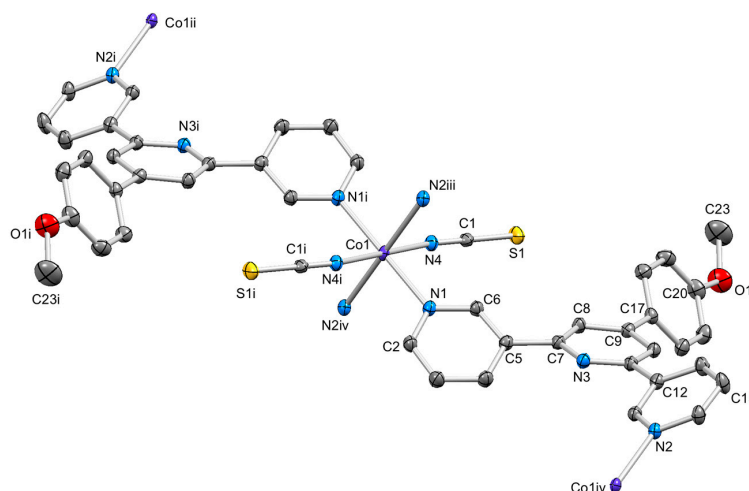
### 3.4. Structures of $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$ and $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$

$[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$  crystallize in the monoclinic space groups  $P2_1/n$  and  $P2_1/c$ , respectively and are pseudopolymorphs. We discuss only the structure of  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  in detail, since the networks in the two structures are essentially the same. Figure 7 shows an overlay of the repeating units (with symmetry generated atoms) in the two structures and angles between the ring planes given in the caption confirm the similarity of the structures. The six-coordinated cobalt atom lies on an inversion centre.

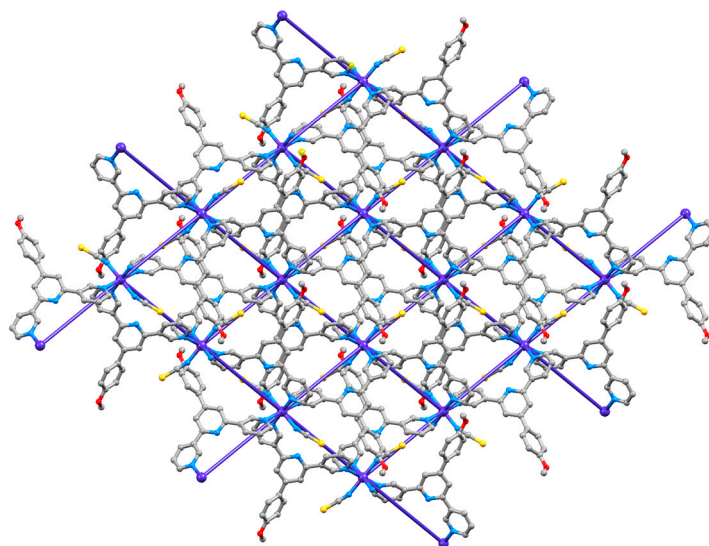
The repeating unit in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  (with symmetry generated atoms) is displayed in Figure 8, and the figure caption gives important bond parameters. All bond distances are as expected. The coordination mode of **2** mimics that in  $[\text{Co}(\mathbf{2})_2(\text{NCS})_2(\text{MeOH})_2]_n$ . Atom Co1 acts as a 4-connecting node and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  is a (4,4) net with the organic ligands acting as linkers (Figure 9). In contrast to the conformation II adopted by **2** in  $[\text{Co}(\mathbf{2})_2(\text{NCS})_2(\text{MeOH})_2]_n$  (see Section 3.2), ligand **2** in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  adopts a conformation similar to Mode III in Scheme 1. The twist angles between the pyridine rings with N1/N3 and N3/N2 are  $32.9^\circ$  and  $32.0^\circ$ , and that between the rings in the 4-methoxyphenylpyridine unit is  $50.8^\circ$ . The latter is significantly greater than the corresponding angle in  $[\text{Co}(\mathbf{2})_2(\text{NCS})_2(\text{MeOH})_2]_n$  ( $12.4^\circ$ ) and reflects the fact that in the 2D-net, ligand **2** does not engage in  $\pi$ -stacking interactions. As Figure 10 shows, the 4-methoxyphenyl units protrude from the upper and lower sides of each (4,4) net and are accommodated in pockets in an adjacent sheet. The closest  $\text{C}-\text{H} \cdots \pi$  contact between sheets is  $3.55 \text{ \AA}$  ( $\text{CH}_{\text{Me}} \cdots \text{centroid}_{\text{pyridine}}$ ) which is rather long to be significant [55].



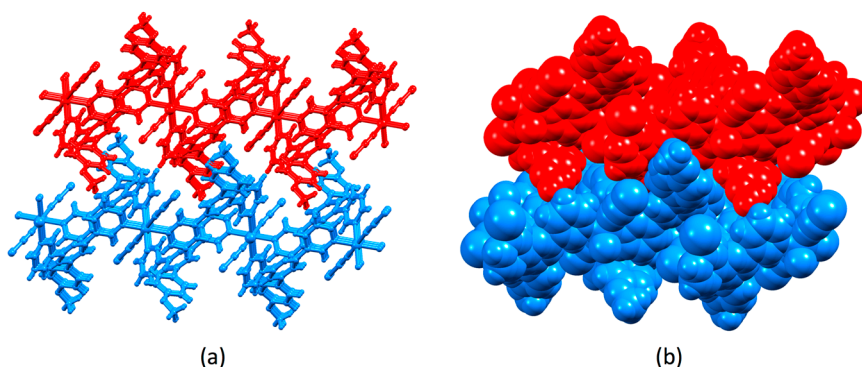
**Figure 7.** Overlay of the repeating units (with symmetry generated atoms) in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  (shown in red) and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$  (in green). Solvent molecules omitted. The angles between the planes of connected pyridine rings are  $32.9^\circ$  and  $32.0^\circ$  in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $33.8^\circ$  and  $38.2^\circ$  in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$ ; the angle between the rings in the 4-methoxyphenylpyridine unit is  $50.8^\circ$  in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $50.5^\circ$  in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$ .



**Figure 8.** The repeating unit in the 2D network in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  with symmetry generated atoms; ellipsoids are plotted at a 40% probability level, and H atoms omitted for clarity. Symmetry codes: i =  $1-x, 1-y, -z$ ; ii =  $1/2-x, 1/2+y, -1/2-z$ ; iii =  $-1/2+x, 1/2-y, -1/2+z$ ; iv =  $3/2-x, 1/2+y, 1/2-z$ . Selected bond distances: Co1–N4 2.0884(16), Co1–N2<sup>iii</sup> = 2.2096(16), Co1–N1 = 2.1714(16), O1–C20 = 1.366(3), O1–C23 = 1.425(4) Å; C20–O1–C23 =  $117.3(2)^\circ$ , by symmetry, N1–Co1–N1<sup>i</sup> = N2<sup>iii</sup>–Co1–N2<sup>iv</sup> = N4–Co1–N4<sup>i</sup> =  $180.0^\circ$ .



**Figure 9.** Part of one 2D network of  $[\text{Co}(\mathbf{2})_2(\text{NCS})_2(\text{MeOH})_2]_n$  showing the 4-connecting Co nodes in the (4,4) net.



**Figure 10.** Parts of two adjacent 2D sheets in  $[\text{Co}(\mathbf{2})_2(\text{NCS})_2(\text{MeOH})_2]_n$ : (a) ball-and-stick and (b) space-filling representations.

In  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$ , SQUEEZE had to be applied to the solvent region, whereas in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$ , the solvent was modelled as one  $\text{CHCl}_3$  molecule disordered around an inversion centre and one  $\text{CHCl}_3$  disordered over 2 orientations. The  $\text{CHCl}_3$  molecules occupy sites both within and between the 2D sheets, and presumably MeOH molecules in  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  occupy analogous sites.

#### 4. Conclusions

We have described the assembly of 1D-coordination polymers or 2D-networks from reactions of  $\text{Co}(\text{NCS})_2$  with ligand **2**. In a series of experiments, crystals were grown at room temperature by layering a MeOH solution of  $\text{Co}(\text{NCS})_2$  over a  $\text{CHCl}_3$  solution of **2**. With a  $\text{Co}(\text{NCS})_2$  to **2** molar ratio of 1:1 or 2:1, the product was  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$  consisting of 1D-chains with coordinated methanol. Powder XRD confirmed that the crystals selected for single-crystal X-ray diffraction or cell dimension checks were representative of the bulk sample. Longer crystallization times with a  $\text{Co}(\text{NCS})_2$ :**2** molar ratio of 1:1 yielded crystals of  $[\text{Co}(\mathbf{2})(\text{NCS})_2(\text{MeOH})_2]_n$  (1D-chain) and the pseudopolymorphs  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 3\text{MeOH}]_n$  and  $[\{\text{Co}(\mathbf{2})_2(\text{NCS})_2\} \cdot 2.2\text{CHCl}_3]_n$  (2D-nets), each type of crystal originating from a different zone in the crystallization tube. PXRD for this last experiment confirms that the dominant product in the bulk sample was the 1D-coordination polymer. This investigation highlights the interplay between thermodynamic and kinetic products from crystallization processes, and illustrates once again [56] that we are far from fully understanding

the assembly algorithms of metallo-coordination polymers involving terpyridine ligands. In particular, these experiments emphasize that, although the crystalline products represent local thermodynamic minima, the growth process is far from equilibrium and the local concentrations and ratios of reactants are time dependent.

**Supplementary Materials:** The following are available online at <http://www.mdpi.com/2073-4360/11/7/1224/s1>, Figure S1 Electrospray mass spectrum of **2**; Figures S2–S5: NMR spectra of **2**; Figures S6–S9: powder X-ray diffraction data.

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